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*Note on Philadelphite—a new mineral.*—MR LEWIS gave a preliminary description of a new vermiculite from near Wayne Station on the Germantown Railroad, which he proposed to call “Philadelphite.” It occurs in plates of a brown color and talcose lustre, existing as seams in an altered hornblende rock. When heated, it exfoliates with great force to many times its original size and becomes of a coppery bronze color. It was stated that while exfoliating, it was able to lift over 50,000 times its own weight. It had a hygroscopic power nearly as great as that of chloride of calcium. Its optical characters and its chemical composition were given.

*Analysis of Philadelphite.*—MR. REUBEN HAINES contributed the following analyses of Philadelphite.

Specific gravity (determined in alcohol of 95 p. c.) 2.78–2.96.

	I.	II.
SiO <sub>2</sub>	39.06	38.52
Fe <sub>2</sub> O <sub>3</sub>	20.59	20.01
Al <sub>2</sub> O <sub>3</sub>	14.75	14.82
Fe <sub>2</sub> O	2.04	2.04
CaO	.99	1.08
MgO	11.49	11.32
MnO (traces)	....	....
Li <sub>2</sub> O (traces)	....	....
K <sub>2</sub> O	6.89	6.61
Na <sub>2</sub> O	.90	.64
H <sub>2</sub> O	4.27	4.27
F (traces)	....	....
	<hr/> 100.98	<hr/> 99.31

Per cent. of hygroscopic water in I, 3.12 p. c.; in II, 3.43 p. c.

In these analyses the mineral was dried at 100° C., the hygroscopic water not being included in the determinations. Owing to its very hygroscopic nature, it was found very difficult to obtain its weight at 100° C. accurately. It gains rapidly in weight while being weighed upon the balance. Examples of its hygroscopic power were given. The analyses were made by dissolving the mineral in concentrated hydrochloric acid. Iron was estimated volumetrically and the alkalies by Smith's method of fusion.